

Ligand exchange in a supersonic expansion

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The smallest possible molecule that involves hydrogen atoms forming a van der Waals complex is dihydrogen itself. Complexes of dihydrogen with other binding partners are complicated by the fact that H₂ can exist in two different nuclear spin isomeric forms, one of which must retain its rotation in the complex (*ortho*-H₂, $I = 1$) whereas the other does not (*para*-H₂, $I = 0$). This results in different physical behaviour and can even impact their reactivity in the gas phase.[1] Hence, seemingly simple complexes are not necessarily easy to understand and it is commonly observed that *ortho*-H₂ complexes are more stable than *para*-H₂.

Inspired by a rotational study on the OCS-H₂ complex[2], we investigated the formation of the *ortho*-H₂, *para*-H₂ and He complex with benzonitrile dependent on the H₂ content in the expansion using cavity-resonator rotational jet spectroscopy. Similar to OCS-H₂, an enhancement of the signal intensity of the *ortho*-complex with increasing H₂ content was observed, whereas an inverse behaviour was found for the *para* and He-complex. To understand this behaviour we kinetically modelled their intensity on the basis of an often proposed ligand exchange mechanism. These results may have implications for the *ortho/para*-H₂ ratio observed in the interstellar medium (ISM) for which polycyclic aromatic hydrocarbons (PAHs) have been found to be important.[3,4] Benzonitrile can serve as a small model system for PAHs and has itself been detected in the ISM.[5]

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[2] Z. Yu et al. J. Chem. Phys. 2007 127, 054305.

[3] E. Bron et al. A&A 2016 588, A27.

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Keywords

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